

REMARKS

In the Final Office Action dated June 21, 2006 the previously made restriction requirement was withdrawn. Accordingly claims drawn to additives containing calcium and magnesium ions, as well as claims to additives containing other polyvalent ions continue to be presented for consideration.

Reorganization of the claims for re-examination

For the purpose of simplifying the presently requested Continuing Examination, the applicant, acting through the undersigned attorney, has canceled the original claims and submitted formal "new" Claims 21-48. To further simplify the task of examination, Claims 21-34 are drawn to the "calcium chloride, magnesium chloride and calcium oxide" subject matter. Claims 35 - 48 are drawn to the subject matter of "polyvalent ions,," now recited as Fe^{++} , Fe^{+++} , Cu^{++} and Ba^{++} .

The presently submitted claims represent editorial changes made for the purpose of placing the claims in better form, and also amendments of substantive subject matter which is fully supported by the original disclosure and avoids all art cited during the prosecution of this application.

The rejection over prior art has been overcome or is inapplicable to the present claims

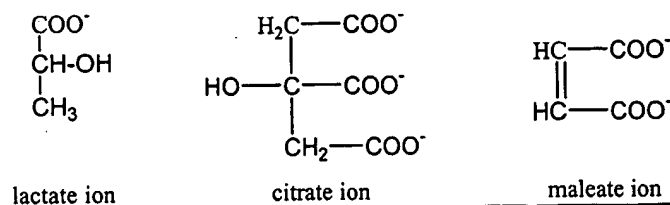
To show the *prima facie* allowability of all of the currently presented claims the applicant, acting through the undersigned attorney, discusses all pertinent references cited for anticipation or obviousness during the prosecution of this application, and shows that the present claims obviate all rejections based on these references.

Claims were rejected in the Office Action dated September 21, 2005 over published United States patent application US 2004/0058054 A1 (*Elder et al.*). This reference discloses

reduction of acrylamide formation in thermally processed foods by adding calcium lactate, calcium citrate or calcium maleate to the food before thermal processing. The application also elaborates on the theory that acrylamide formation normally occurs because of decomposition or reaction of the amino acid asparagine in the food. Thus, the adding of calcium lactate, calcium citrate or calcium maleate to the food minimizes decomposition of asparagines to acrylamide during a thermal process.

Calcium lactate is only sparingly soluble in water. According to the enclosed *Handbook of Chemistry and Physics* 51st Edition (1970 – 1971) pp. B-77, 78, calcium lactate's solubility is 3.1g in 100 cc water. According to the same source, calcium citrate is even less soluble (0.85g in 100 cc of water) and calcium maleate has a solubility of 2.89g in 100 cc of water. Given the fact that in doughs, baked goods and the like the water content is relatively modest, it is unlikely that calcium lactate, calcium citrate or calcium maleate contribute a significant amount of calcium ions (Ca^{++}) to the result described in the *Elder et al.* reference.

Even more importantly, the lactates, citrates and maleates or organic salts, each have the respective structural formulas shown below.



Because the *Elder et al.* reference teaches that the decomposition or reaction of asparagines results in acrylamide formation and that addition of these organic salts minimizes such decomposition or reaction, a person of ordinary skill in the art is likely to conclude that it is the

“organic” portion of these molecules which react or interact with asparagines and minimize the formation of acrylamide.

The present Claims 21-34 do not encompass the use of calcium lactate, citrate or maleate. Therefore any rejection for anticipation by *Elder et al.* is clearly avoided. For the reasons explained above a person of ordinary skill would not find it obvious to use calcium chloride, magnesium chloride or calcium oxide to minimize the formation of acrylamide. For this reason the subject matter of Claims 21 – 34 is not obvious over the *Elder et al.* reference or over any other reference of record.

“To serve as an anticipation when the reference is silent about the asserted inherent characteristic, such gap in the reference may be filled with recourse to extrinsic evidence. Such evidence must make clear that the missing descriptive matter is necessarily present in the thing described in the reference, and that it would be so recognized by persons of ordinary skill.”

Continental Can Company USA, Inc. v. Monsanto Co., 948 F.2d 1264, 20 USPQ2d 1746 (Fed. Cir. 1991).

United States Patent No. 6,436,458 (*Kuechle et al.*) was cited in the Office Action of September 21, 2005. Applicant respectfully submits that this reference actually teaches away from the present claims. First, the main concern of *Kuechle et al.* is to provide a scoopable dough. The reference is not at all concerned and does not describe elimination or minimization of acrylamide formation during the cooking, baking or frying process.

Even more importantly, the reference contains several passages indicating that the presence of iron ions and other ions are to be avoided in the scoopable dough. Column 2, lines 6-9 of the reference state that the dough can include “flour enrichment with all reduced iron, a binder of metal ions, an organic acid and the like.” (underlining added). Column 4, lines 5-6

state that the scoopable dough “can include flour enrichments having iron that is substantially all reduced” (underlining added).

Column 10, lines 13 through 21 state:

“A scoopable dough of the invention can also include a binder of metal ions. Some metal ions can help catalyze enzymatic reactions that can result in dough graying. Thus, a binder of metal ions can be effective against a scoopable dough of the invention losing its natural color. Examples of binders of metal ions include metal chelators and organic acids. A metal chelator can include, for example, ethylene diaminetetracetate (EDTA).” (underlining added).

The term “all reduced iron” necessarily refers to metallic iron and not to iron ions. As is well known in the art “a binder of metal ion” is a chemical that removes the metal ion from effectiveness as an ion, by chelating it. EDTA is a very well known chelator of many kinds of metal ions, including that of iron ions and aluminum ions. Therefore, even if some of the “all reduced iron” in the *Kuechle et al.* disclosure were still in the ionic stage (two valent ferrous instead of three valent ferric), such ion would be removed by chelation. For all these reasons, the *Kuechle et al.* reference teaches away from the subject matter of the present Claims 35-48 and neither anticipates these claims nor renders them obvious.

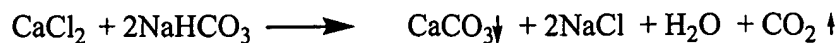
United States Patent No. 5,556,655 (*Vadlamani et al.*) was cited in the Final Office Action of March 7, 2006. This reference is not concerned with eliminating or reducing the formation of acrylamide in foods which are subjected to thermal processes, such as cooking, baking or frying. Rather, its purpose is to reduce darkening of the food product during storage. The majority of examples of this reference are directed to a gluten product. The salts or ions added to the raw product (before it may be heated such as baked) were zinc and aluminum. The use of zinc and/or aluminum ions is no longer recited in the presently amended claims. Because

the reference is entirely silent about reducing acrylamide formation, a person of ordinary skill in the art would receive no hint or suggestion that the use of Fe^{++} , Fe^{+++} , Cu^{++} and/or Ba^{++} ions, as stated in Claims 35 – 48, will reduce formation of acrylamides.

United States Patent No. 5,882,712 (*Wu*), cited in the Office Action of June 21, 2006, in spite of its appearance at first glance, does not disclose the use or presence of Ca *ions* originating from CaCl_2 , CaO or from other source of calcium, or of magnesium ions originating from MgCl_2 or from other source of magnesium while the therein disclosed dough or like batter product is baked or otherwise subjected to heat. On the contrary, this reference teaches to one of ordinary skill in the art that calcium or magnesium is no longer present in an ionic form, nor as CaCl_2 or MgCl_2 after the leavening process has been completed and the dough or batter is baked. The reasons are as follows.

The *Wu* reference discloses the use of a leavening composition for leavening a dough or batter before the dough or batter is baked or otherwise subjected to heat. The leavening composition of the reference includes a “leavening acid” and a “basic compound” which when reacted with the acid generates carbon dioxide. Carbon dioxide (CO_2) is a gas that forms bubbles in the dough and therefore causes it to rise. This is well known in the art. It is significant that the reference refers to CaCl_2 or MgCl_2 as the “acid.” In the ensuing discussion CaCl_2 will be used as an example although the discussion is equally applicable to MgCl_2 .

As a person of skill in the art will readily understand that carbon dioxide (CO_2) gas forms in the reaction between CaCl_2 and the carbonate containing “basic” compound (such as sodium bicarbonate, NaHCO_3) according to the following chemical equation.



As is well known in the art CO₂ is the gas that causes leavening of the batter, and calcium carbonate (CaCO₃) is water insoluble (a substance that actually forms many mountain chains on the Earth). Water insoluble calcium carbonate does not contribute calcium ions (Ca⁺⁺) to the batter like water soluble calcium chloride (CaCl₂) does. The same type of chemical equation and discussion is applicable to water soluble magnesium chloride (MgCl₂) and water insoluble MgCO₃.

Because CaCl₂ and MgCl₂ is initially added to the batter or dough to leaven together with a basic compound (such as NaHCO₃) and because, as shown above, calcium or magnesium ions are no longer present in substantive amounts in the batter or dough after it is leavened and while it is baked, the added calcium or magnesium of this reference does not inherently reduce acrylamide formation. Nor would a person of ordinary skill in the art find it obvious to omit the "basic compound" (such as NaHCO₃) from the leavening substance added to the dough for the purpose of reducing acrylamide formation while the dough is baked.

The presently submitted independent Claims 21, 32 and 33 clearly set forth that the additive contributes calcium or magnesium ions to the dough or batter to be present while the dough or batter is cooked under heat. Therefore, the subject matter of Claims 21 – 34 clearly avoids the *Wu* reference.

United States Patent No. 4,272,554 (*Schroeder et al.*) cited in the Final Office Action of June 21, 2006 discloses preparing blister-inhibited potato chips. The disclosed process involves the step of applying small amount of calcium (a calcium chloride or other calcium salt) to the

surface of the chips by "spraying, dipping or the like." Present claims 21-34 are not drawn to the making of fried or baked potato chips. This disclosure does not suggest to one of ordinary skill in the art that formation of acrylamide is reduced when calcium ions are added to a dough or batter. Therefore, the present claims clearly avoid the *Schroeder et al.* reference.

In light of the foregoing, all claims are free of the prior art and are in *prima facie* allowable condition. Their early allowance is respectfully solicited.

In the event the Examiner is of the opinion that a telephone conference with the undersigned attorney would materially facilitate the final disposition of this case, she is respectfully requested to telephone the undersigned attorney at the below listed telephone number.

I hereby certify that this correspondence is being deposited with the U.S. Postal Service as first class mail in an envelope addressed to: Commissioner for Patent, PO Box 1450, Alexandria, VA 22313-1450 on August 31, 2006.


Very truly yours,

SNELL & WILMER L.L.P.

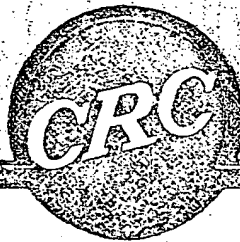
By: Sharon Farnus

Signature

Dated: August 31, 2006



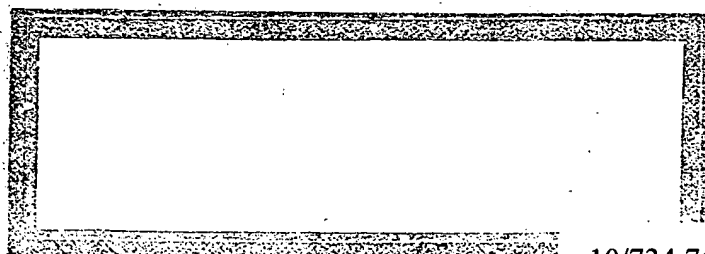
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HANDBOOK
of
CHEMISTRY
and
PHYSICS

51ST

EDITION
1970 - 1971



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10/734,766
Yoshi Tomoda et al.
Attachment to Response
Page 1 of 3

PHYSICAL CONSTANTS OF INORGANIC COMPOUNDS (Continued)

No.	Name	Synonyms and Formulae	Mol. wt.	Crystalline form, properties and index of refraction	Density or spec. gravity	Melting point, °C	Boiling point, °C	Solubility, in grams per 100 cc			No.
								Cold water	Hot water	Other solvents	
Calcium											
c129	iodide, hexahydrate	CaI ₂ ·6H ₂ O	401.98	yel, hex need	2.55	d 42	160	757 ^a	1680 ^a	s a, al, acet	c174
c130	iron (III) aluminate	Calcium (tetra-) aluminato-ferite, nat. celite. 4CaO·Fe ₂ O ₃ ·Al ₂ O ₃	485.97	brn, rhomb, 1.98, 2.05, 2.08 all for λ	3.77	1418 _g		3			c175
c131	isobutyrate	Ca(C ₄ H ₇ O ₂) ₂ ·5H ₂ O	304.35	col powd				20	al s		c176
c132	lactate	Ca(C ₃ H ₅ O ₂) ₂ ·5H ₂ O	308.30	wh need, eff		-3H ₂ O, 100		3.1 ^a	7.9 ^a	al s a; i al, eth	c177
c133	laurate	Ca(C ₁₁ H ₂₁ O ₂) ₂ ·H ₂ O	458.73	wh need, eff		182-183		0.004 ^u	0.055 ^u	0.059 ^u , 1.72 ^u al	c178
c134	linoleate	Ca(C ₁₈ H ₃₃ O ₂) ₂	598.97	wh amorph powd				i		s al, eth	c179
c135	magnesium carbonate	Nat. dolomite. CaCO ₃ ·MgCO ₃	184.41	col, trig, 1.6817, 1.5026	2.872	d 730-760		0.032 ^u			c180
c136	magnesium metasilicate	Nat. diopside. CaO·MgO·2SiO ₂	216.56	col, monocl, 1.665, 1.672, 1.695	3.275	1391		i	i	i HCl	c181
c137	magnesium orthosilicate	Nat. mervinit. 3CaO·MgO·SiO ₂	328.72	col to pa grn, monocl, 1.708, 1.711, 1.718	3.150						c182
c138	di-malate	CaC ₂ H ₃ O ₄ ·3H ₂ O	226.20	col, rhomb, 1.545, 1.555, 1.575				0.321 ^a	0.451 ^a	i al	c183
c139	l-malate	CaC ₂ H ₃ O ₄ ·2H ₂ O	208.18	col.				0.812 ^a	1.224 ^a	s al	c184
c140	malate, dihydrogen	Ca(HC ₂ H ₃ O ₄) ₂ ·6H ₂ O	414.33	rhomb, or wh cr powd, 1.493, 1.507, 1.545				al s			c185
c141	maleate	CaC ₂ H ₂ O ₄ ·H ₂ O	172.15	col rhomb, 1.495, 1.575, 1.640				2.89 ^u	3.21 ^u		c186
c142	malonate	CaC ₂ H ₂ O ₄ ·4H ₂ O	214.19					0.44 ^a	0.72 ^u		c187
c143	permanganate	Ca(MnO ₄) ₂ ·5H ₂ O	368.03	purp cr	2.4	d		331 ^u	338 ^u	s NH ₄ OH	c188
c144	α-methylbutyrate	Calcium ethylmethylacetate. Ca(C ₄ H ₉ O ₂) ₂	242.34					24.24 ^a	25.65 ^u		c189
c145	molybdate	Nat. pawellite. CaMoO ₄	200.01	col, tetr, 1.967, 1.978	4.38-4.53			i	d	s a i al, eth	c190
c146	nitrate	Ca(NO ₃) ₂	164.09	col, cub, hyg	2.504 ^u	561		121.2 ^u	376 ^u	14 ^u al; s MeOH; liq NH ₃ , acet; i eth	c191
c147	nitrate, tetrahydrate	Ca(NO ₃) ₂ ·4H ₂ O	236.15	col, monocl, deliq, 1.465, 1.498, 1.504	α 1.896, β 1.82	α 42.7, β 39.7	d 132	266 ^a	660 ^a	s al, acet	c192
c148	nitrate, trihydrate	Ca(NO ₃) ₂ ·3H ₂ O	218.14	col, tricl.		51.1		d	d	s dil a; d abs al	c193
c149	nitride	Ca ₃ N ₂	148.25	brn cr, hex	2.63 ^u	1195		45.9 ^a	89.6 ^u	al s al	c194
c150	nitrite	Ca(NO ₂) ₂ ·H ₂ O	150.11	col-yelsh, hex, deliq	2.23 ^u	-H ₂ O, 100					c195
c151	nitrite, tetrahydrate	Ca(NO ₂) ₂ ·4H ₂ O	204.15	col cr, tetr.	1.674 ^a	-2H ₂ O, 44		74.9 ^a	106 ^u	s al	c196
c152	oleate	Ca(C ₁₈ H ₃₃ O ₂) ₂	603.01	wh wax-like cr		83-84		0.04 ^u	0.03 ^u	al s eth	c197
c153	oxalate	CaC ₂ O ₄	128.10	col, cub	2.2 ^a	d		0.00067 ^u	0.0014 ^u	al s i ac a	c198
c154	oxalate, hydrate	CaC ₂ O ₄ ·H ₂ O	146.12	col.	2.2	-H ₂ O, 200		i	i	s a; i ac a	c199
c155	oxide	Lime, calcia. CaO	56.08	col, cub, 1.838	3.25-3.38	2580	2850	0.131 ^u d	0.07 ^u d	s a	c200
c156	oxide, per	CaO ₂	72.08	wh, tetr, 1.895	2.92 ^u	d 275		al s		s a	c201
c157	oxide, per-octahydrate	CaO ₂ ·8H ₂ O	216.20	wh, tetr, pearly	1.70	-8H ₂ O, 200	d 275 expl	al s	d	s a, NH ₄ salts; i al, eth	c202
c158	palmitate	Ca(C ₁₆ H ₃₁ O ₂) ₂	550.93	wh or yelsh, wh fatty powd				0.003 ^u		v al s al; 0.008 ^u eth	c203
c159	l-phenol-4 sulfonate(p-)	Ca[C ₆ H ₄ (OH)SO ₃] ₂ ·H ₂ O	404.43	wh to pinkish powd						s al	c204
c160	phenoxide	Ca(OC ₆ H ₅) ₂	226.29	redsh powd				al s		al s al	c205
c161	hypophosphate	Ca ₂ P ₂ O ₅ ·2H ₂ O	274.13	gel				i		s HCl	c206
c162	metaphosphate	Ca(PO ₃) ₂	198.02	col, 1.588, 1.595	2.82	975		i	i	al	c207
c163	orthophosphate, di-(sec)	Nat. brushite. CaH ₂ P ₂ O ₇ ·2H ₂ O	172.09	wh, tricl, 1.5576, 1.5457, 1.5392	2.306 ^u	-H ₂ O, 109		0.0316 ^u	0.075 ^u	i al, s a	c208
c164	orthophosphate, mono-(prim.)	Ca(H ₂ PO ₄) ₂ ·H ₂ O	252.07	col, tricl, deliq, 1.5292, 1.5176, 1.4392	2.220 ^u	-H ₂ O, 109	d 203	1.8 ^u	d	s a	c209
c165	orthophosphate, tri-(tert.)	Nat. whitlockite. Ca ₃ (PO ₄) ₂	310.18	wh amorph powd, 1.629, 1.626	3.14	1670		0.002	d	i al; s a	c210
c166	pyrophosphate	Ca ₂ P ₂ O ₇	254.10	col, biax, 1.585, 1.604	3.09	1230		i		s a	c211
c167	pyrophosphate, pentahydrate	Ca ₂ P ₂ O ₇ ·5H ₂ O	344.18	col, monocl, 1.539, 1.545, 1.551	2.25			al s		s a; i NH ₄ Cl	c212
c168	phosphide	Ca ₃ P ₂	182.19	gray lumps	2.51	ca 1600		d ev PH ₃		s a; i al, eth, bs	c213
c169	hypophosphite	Ca(H ₂ PO ₂) ₂	170.06	wh-gray, monocl		d		15.4 ^u	12.5 ^u	i al	c214
c170	orthophosphite, di-	2CaHPO ₃ ·3H ₂ O	294.17					al s	d	s NH ₄ Cl	c215
c171	orthoplumbate	Ca ₂ PbO ₄	351.35	red-br cr	5.71	d		i	d	s a	c216
c172	propionate	Ca(C ₃ H ₇ O ₂) ₂ ·H ₂ O	201.24	col, monocl tabl.				49 ^a	55.8 ^u	i al	c217
c173	l-quinate	Ca(C ₇ H ₁₅ O ₅) ₂ ·10H ₂ O	602.56	rhomb leaf		50, -10H ₂ O 120		16 ^u		i al	c218

PHYSICAL CONSTANTS OF INORGANIC COMPOUNDS (Continued)

No.	Name	Synonyms and Formulae	Mol. wt.	Crystalline form, properties and index of refraction	Density or spec. gravity	Melting point, °C	Boiling point, °C	Solubility, in grams per 100 cc		
								Cold water	Hot water	Other solvents
c87	Calcium butyrate	$\text{Ca}(\text{C}_4\text{H}_7\text{O}_2)_2 \cdot 3\text{H}_2\text{O}$	268.32	col cr.				s	sl s	
c88	carbide	CaC_2	64.10	col, tetr, 1.75	2.22	stab 25-447	2300	d	d	
c89	carbonate	Nat. aragonite. CaCO_3	100.09	col, rhomb, 1.530, 1.681, 1.685	2.930	tr to calcite 520	d 825	0.00153 ²⁵	0.00190 ²⁵	s a, NH_4Cl
c90	carbonate	Nat. calcite. CaCO_3	100.09	col, rhomb or hex, 1.6583, 1.4864	2.710 ¹⁵	1339 ¹⁰²⁵	d 898.6	0.0014 ²⁵	0.0018 ²⁵	s a, NH_4Cl
c91	carbonate, hexahydrate	$\text{CaCO}_3 \cdot 6\text{H}_2\text{O}$	208.18	col, monocel, 1.460, 1.535, 1.545	1.771 ^o					
c92	chlorate	$\text{Ca}(\text{ClO}_3)_2$	206.99	wh cr, hyg.		340 ± 10 (-some O)		s	s	s al, acet
c93	chlorate, dihydrate	$\text{Ca}(\text{ClO}_3)_2 \cdot 2\text{H}_2\text{O}$	243.01	wh-yelsh, rhomb, or monocel, deliq	2.711	-H ₂ O, 100	177.7 ¹	v s	v s	s al, acet
c94	perchlorate	$(\text{CaClO}_4)_2$	238.98	col cr.	2.651	d 270		188.6 ²⁵	v s	166.2 ²⁵ al; 237.4 MeOH
c95	chloride	CaCl_2	110.99	col, cub, deliq 1.52	2.15 ²⁵	772	>1600	74.5 ²⁰	159 ¹⁰⁰	s al, acet, ace a
c96	chloride aluminate	$3\text{CaO} \cdot \text{Al}_2\text{O}_3 \cdot \text{CaCl}_2 \cdot 10\text{H}_2\text{O}$	561.33	col, monocel or hex, 1.550, 1.535	1.892 ¹⁴	-H ₂ O, 105	-8H ₂ O, 350	sl s	d	s a
c97	chloride, dihydrate	$\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$	147.02	col cr.	0.835			97.7 ^o	326 ⁶⁰	50 ⁹⁰ al
c98	chloride, hexahydrate	$\text{CaCl}_2 \cdot 6\text{H}_2\text{O}$	219.08	col, trig, deliq, 1.417, 1.393	1.71 ²⁵	29.92	-4H ₂ O, 30, -6H ₂ O, 200	279 ⁹	536 ²⁰	s al
c99	chloride, monohydrate	$\text{CaCl}_2 \cdot \text{H}_2\text{O}$	129.00	col cr, deliq.		260		76.8 ^o	249 ¹⁰⁰	s al; i acet
c100	chloride fluoride orthophosphate	$3\text{Ca}_2(\text{PO}_4)_2 \cdot \text{CaCl}_2 \cdot \text{CaF}_2$	1025.08	col cr. 1.634, 1.631	3.14	1270		v sl s		
c101	chlorite	$\text{Ca}(\text{ClO}_2)_2$	174.98	wh, cub	2.71			d	d	i al
c102	hypochlorite	$\text{Ca}(\text{ClO})_2$	142.98	wh powd or flat pl. 1.545, 1.69	2.35	d 100		s		i al
c103	chlorite, basic	$\text{Ca}(\text{ClO})_2 \cdot 2\text{Ca}(\text{OH})_2$	257.16	wh, hex, 1.51, 1.585	2.10			sl s solns with 5-6 % avail Cl	d	d a
c104	Hypochlorite, basic	Bleaching powder, chlorinated lime. $\text{Ca}(\text{ClO})_2 \cdot \text{CaCl}_2 \cdot x\text{Ca}(\text{OH})_2 \cdot x\text{H}_2\text{O}$	comp varies	wh powd strong Cl odor		d		d evln Cl ₂		d a
c105	Hypochlorite, trihydrate	$\text{Ca}(\text{ClO})_2 \cdot 3\text{H}_2\text{O}$	197.03	tetr pl. 1.535, 1.63	2.1	-3H ₂ O, 60				
c106	chromate	$\text{CaCrO}_4 \cdot 2\text{H}_2\text{O}$	192.09	yel, monocel pr.		-2H ₂ O, 200		16.3 ²⁰	18.2 ¹⁵	s a, al
c107	chromite	CaCr_2O_7	208.07	ol grn, cub need.	4.8 ¹⁵	2090		i	i	s a; s fus K_2CO_3
c108	cinnamate	$\text{Ca}(\text{C}_6\text{H}_5\text{O}_2)_2 \cdot 3\text{H}_2\text{O}$	388.44	col cr.				0.22 ¹	1.34 ¹⁰⁰	
c109	citrate	$\text{Ca}_3(\text{C}_6\text{H}_5\text{O}_7)_2 \cdot 4\text{H}_2\text{O}$	570.51	wh need.		-4H ₂ O, 120		0.85 ¹⁵	0.96 ²⁵	0.0065 ¹⁵ al
c110	cyanamide	CaCN_2	80.10	col, hex, rhbdr.		1300 subl >1150		d evl NH_3	d	
c111	cyanide	$\text{Ca}(\text{CN})_2$	92.12	wh powd.		d >350		d	d	
c112	cyanoplatinite	$\text{CaPt}(\text{CN})_4 \cdot 5\text{H}_2\text{O}$	429.31	yel-grn fluoresc, rhomb, 1.6226		-5H ₂ O, 100		s		
c113	ferricyanide	$\text{Ca}_3[\text{Fe}(\text{CN})_6]_2 \cdot 12\text{H}_2\text{O}$	760.42	red need, deliq				v s	v s	
c114	ferrite, mono	$\text{CaO} \cdot \text{Fe}_2\text{O}_3$	215.77	dk redsh r, rhomb, 2.58, 2.43 (Na)	5.08	1250		i	i	v sl s a
c115	ferrocyanide	$\text{Ca}_2\text{Fe}(\text{CN})_{11}$ or $12\text{H}_2\text{O}$	490.28	yel triel, 1.570, 1.582, 1.596	1.68	d		86.8 ²⁵	115 ⁶⁴	i al
c116	fluosilicate	CaSiF_6	182.16	col, tetr.	2.66 ¹⁵			sl s		s al, HF, HCl
c117	fluoride	Nat. fluorite. CaF_2	78.08	col, cub luminisc w heat, 1.434	3.180	1340	ca 2500	0.0016 ¹⁵	0.0017 ²⁵	s NH_4 salts; sl s a; i acet
c118	fluosilicate, dihydrate	$\text{CaSiF}_6 \cdot 2\text{H}_2\text{O}$	218.19	col, tetrag.	2.254			sl s d		s HCl, HF; i al
c119	formate	$\text{Ca}(\text{CHO}_2)_2$	130.12	col, rhomb, 1.510, 1.514, 1.578	2.015	d		18.2 ^o	18.4 ¹⁰⁰	i al
c120	fumarate	$\text{CaC}_4\text{H}_2\text{O}_6 \cdot 3\text{H}_2\text{O}$	208.18	col, rhomb.				2.11 ²⁰		
c121	d-glucosate	$\text{Ca}(\text{C}_6\text{H}_{11}\text{O}_7)_2 \cdot \text{H}_2\text{O}$	448.40	wh cr powd, need		-H ₂ O, 120		3.3 ¹⁵		v sl s al
c122	glycerophosphate	$\text{CaC}_3\text{H}_5(\text{OH})_2\text{PO}_4$	210.16	wh cr powd, hyg.		d 170		2 ²⁵	sl s	i al
c123	hydride	CaH_2	42.10	wh, rhomb cr.	1.9	816 (in H ₂) d ca 600		d H ₂ + $\text{Ca}(\text{OH})_2$		d a
c124	hydroxide	$\text{Ca}(\text{OH})_2$	74.09	col, hex, 1.574, 1.545	2.24	-H ₂ O, 580	d	0.185 ^o	0.077 ¹⁰⁰	s NH_4 salts, a; i al
c125	hyponitrite	$\text{CaN}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$	172.15	wh cr.	1.834	d 320				d dil a
c126	iodate	Nat. lautarite. $\text{Ca}(\text{IO}_3)_2$	389.89	col, monocel.	4.519 ¹⁵	d 540		0.20 ¹⁵	0.67 ²⁰	s HNO_3 ; i al
c127	iodate, hexahydrate	$\text{Ca}(\text{IO}_3)_2 \cdot 6\text{H}_2\text{O}$	497.98	col, rhomb.		d 35		0.13 ^o	1.22 ¹⁰⁰	s HNO_3
c128	iodide	CaI_2	293.89	yelsh-wh, hex, deliq	3.956 ²⁵	740	ca 1100	209 ²⁰	426 ¹⁰⁰	128 ²⁰ MeOH; s al, acet, a

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Attachment to Response

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